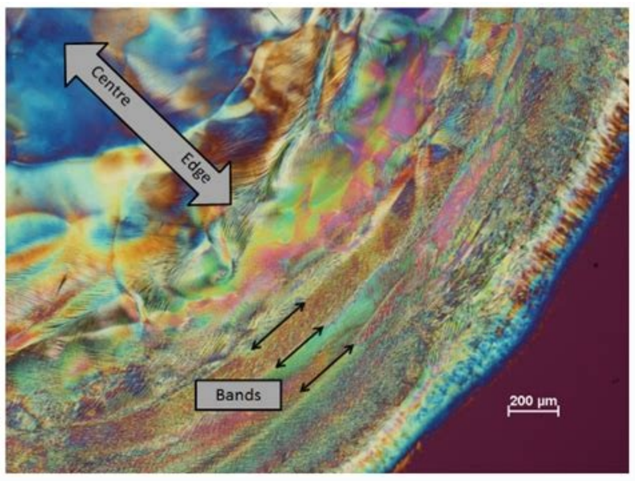
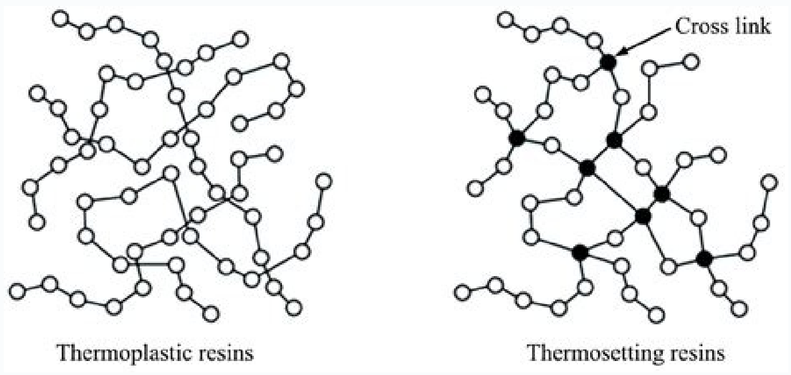
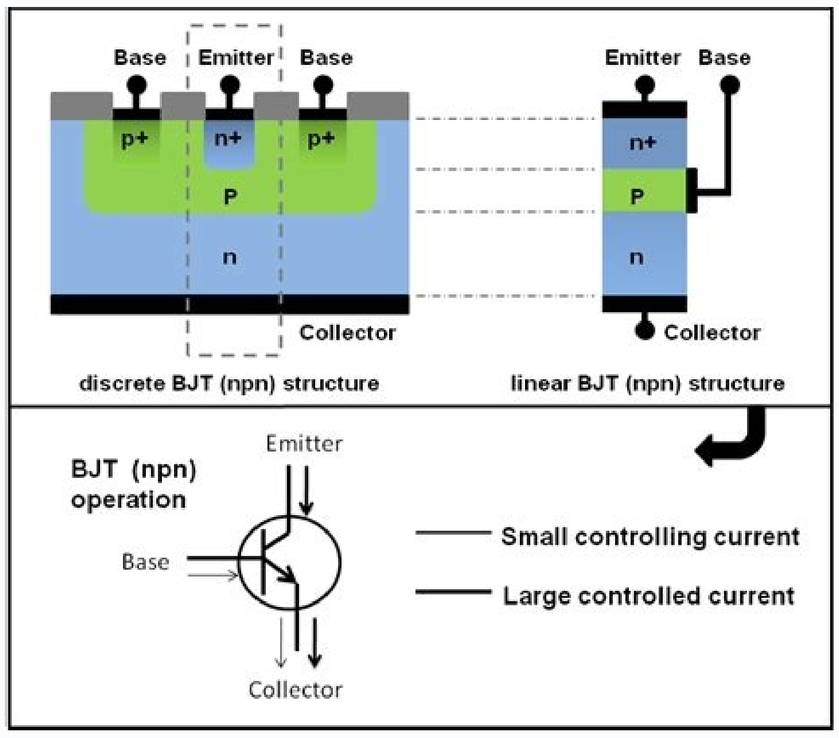
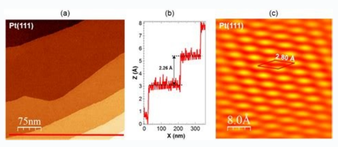
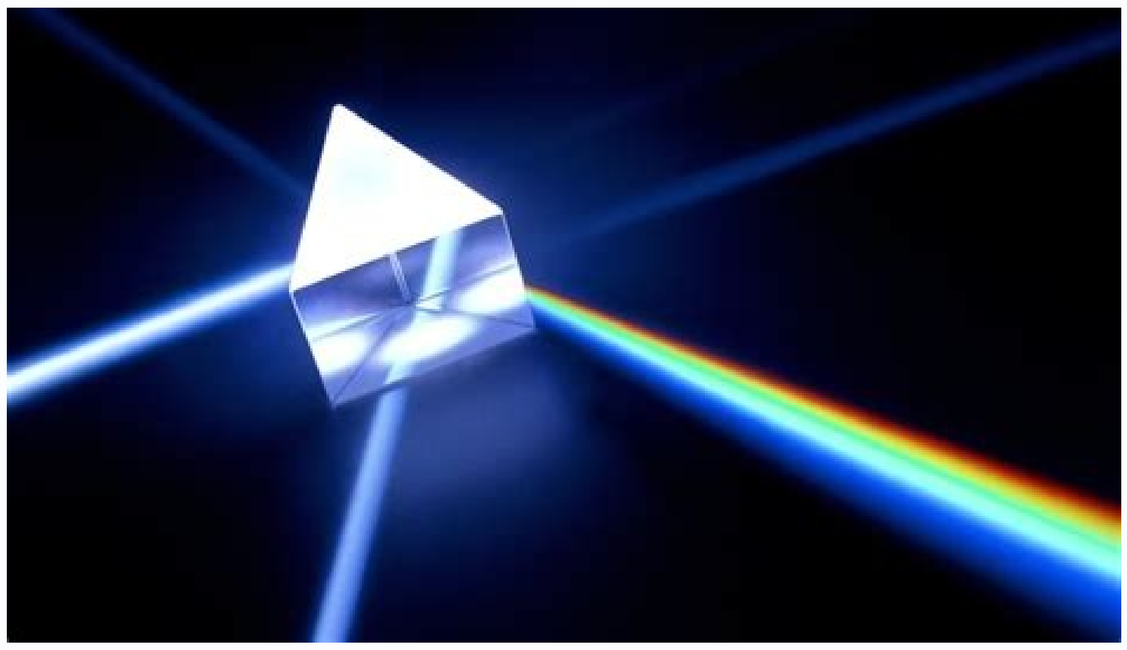
Atomic structure mcq pdf











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The first chapter in Chemistry is the atomic structure, and it builds the foundation of understanding the subject. You should be well versed in the fundamentals of elements and molecules. In this regard, the atomic structure questions will further help in developing a strong understanding of the chapter. Herein, you will find MCQ on the atomic structure chapter so that you can test your knowledge. Besides, you will be prepared to face the exam papers if you are aware of similar questions beforehand. This will boost your confidence in cracking the exam with better grades as well.About Structure of Atom:The 5 Parts of an AtomNucleus Features following :Goldstein in 1886 proposed that the charge to mass rate of the positive patches depends on the nature of the gas which is present in the discharge tube. This means that the charge to mass rate (e/ m) was different for different feasts. He observed that the charge to mass rate of the positive shafts was loftiest in the case of the hydrogen gas that was used in the discharge tube. This is substantial because hydrogen is the lightest atom so m will be the least hence e/ m rate will be loftiest in this case. The flyspeck in the positive shafts in the discharge tube was named a proton. What is a Hydrogen Ion?The hydrogen nucleus is made up of a flyspeck carrying a unit positive electric charge, called a proton. The insulated hydrogen ion, represented by the symbol H+, is thus customarily used to represent a proton. Because the bare nucleus can readily combine with other patches (electrons, tittles, and motes), the insulated hydrogen ion can live only in a nearly flyspeck-free space (high vacuum) and in the gassy state. MCQ on Chemistry - Atomic StructureEvery molecule consists of atoms, and everything we see around us is made up of molecules. Atoms further consist of smaller charged particles that are either positively or negatively charged.The positively charged particles are called protons, and the negative ones are called electrons. An atom also has a neutral particle that is called a neutron. Chemical bonding of various atoms takes place due to the sharing of electrons. Herein, NEET chemistry MCQ on the structure of atom deals with various questions that cover the fundamentals as well as the advanced concepts.1. Who discovered electrons?Rutherford.J. J. Thomson.Neils Bohr.James Chadwick.2. What is the formula for a mass number of an atom?Number of protons + number of electrons.Number of neutrons + number of electrons.Number of protons + number of neutrons.None of these.3. An atom has mass number 37 and atomic number 17. Find the number of protons.4. Which of the following is responsible for the mass of an atom?Only protons.Only neutrons.Neutrons and protons.Protons and electrons.5. Which of the following determines the atomic number of an atom?Number of electrons.Number of protons.Number of electrons and protons.Number of protons and neutrons.6. Nickel has atomic number 28. Choose the correct electronic configuration for Nickel.1s2 2s2 2p4 3s2 3p8 3d101s2 2s2 2p6 3s2 3p6 3d8 4s21s2 2s2 2p4 3s2 3p6 4s21s2 2s2 3s2 3p8 3d107. If both the K and L shells are full, what would be the atomic number of that element?8. Which of the following determines the chemical properties?Number of protons.Number of electrons.Number of neutrons.None of these.9. Atoms that have the same mass number and different atomic number are called?Isotopes.Isotones.Isobars.Isomers. 10. What happens to the atomic number during a chemical reaction?It increases.It changes.Remains the same.Changes alternatively.11. Who proposed the atomic theory?John Dalton.Robert Millikan.J. J. Thomson.Neils Bohr.12. The electronic configuration of atomic number 20 of an atom is which of the following?2, 6, 6, 22, 8, 8, 22, 4, 6, 22, 4, 6, 213. Does an atom differ from an ion with respect to which of the following?Number of protons.Nuclear charge.Number of electrons.Mass number.14. The maximum mass of an atom is concentrated in which of these?Nucleus.Neutrons.Protons.Electrons. 15. The isotopes of a neutral atom of an element differ in which of these?Physical properties.Chemical properties.Atomic number.Mass number.16. What is the sequence to label the subshells in an atom?S, p, d, f, gS, p, p, f, dS, s, p, p, d, f, gS, p, g, d, f17. The electronic configuration for oxygen is written as 1s2 2s2 2p4. Which rule will this configuration be violating?Aufbau’s principle.Hund’s rule.Pauli’s exclusion principle.None of the above.18. If an atom has four unpaired electrons. What is likely to be the total spin of the electron?Answers: 1(b), 2(c), 3(c), 4(c), 5(b), 6(b), 7(c), 8(b), 9(a), 10(c), 11(b), 12(b), 13(c), 14(a), 15(c), 16(a), 17(d), 18(c)Now you are aware of a considerable portion of NEET chemistry, these questions will help you in getting an overall idea of how the multiple-choice questions are framed. A crucial aspect of such questions is that you will often find the options to be similar or close to each other in terms of meaning.Even the slightest misunderstanding can lead you to mark the answer wrong. Therefore, make sure you go through the structure of atom class 11 notes for NEET meticulously. You can not only develop an understanding but also avoid making silly mistakes in the exam.Atomic structure is the fundamental chapter of studying Chemistry. It is thus vital for you to develop a stronghold on this chapter so that you ace the subject as a whole. The MCQ on the atomic structure is meant to assess your understanding and grasp over the minute details included in the chapter.Therefore, make sure you revise the chapter and atomic structure questions thoroughly before solving these. It will help you to judge your level of exam preparation. Also, list down the mistakes and learn from them to not repeat them in your actual exam. Inorganic Chemistry MCQ on Atomic Structure 1. What is the increasing order for the values of e/m for (a) e, p, n, α (b) n, p, e, α (c) n, p, α, e (d) n, α, p, e Answer - (d) n, α, p, e 2. Mass of an atom is equals to which of the following (a)Only protons. (b)Only neutrons. (c)Neutrons and protons. (d)Protons and electrons. Answer -(c)Neutrons and protons. 3. Atomic number of an atom will be (a)Number of electrons. (b)Number of protons. (c)Number of electrons and protons. (d)Number of protons and neutrons. Answer -(b)Number of protons. 4. Nickel has atomic number 28. The correct electronic configuration is (a)1s2 2s2 2p4 3s2 3p8 3d10 (b)1s2 2s2 2p6 3s2 3p6 3d8 4s2 (c)1s2 2s2 2p4 3s2 3p6 4s2 (d)1s2 2s2 3s2 3p8 3d10 Answer -(b)1s2 2s2 2p6 3s2 3p6 3d8 4s2 5. What would be the atomic number of that element having full K and L shell (a)20 (b)14 (c)10 (d)16 Answer -(c)10 6. The chemical properties of an atom related to (a)Number of protons. (b)Number of electrons. (c)Number of neutrons. (d)None of these. Answer -(b)Number of electrons. 7. Atoms having the same mass number and different atomic number are known as (a)Isotopes. (b)Isotones. (c)Isobars. (d)Isomers. Answer -(a)Isotopes. 8. Which of the following quantum number helps to distinguish the electrons of the same orbitals ? (a) Principal quantum number (b) Azimuthal quantum number (c) Magnetic quantum number (d) Spin quantum number Answer -(d) Spin quantum number 9. The maximum electrons accommodate by a sub-shell with n = 6 , l = 2 is (a) 12 electrons (b) 10 electrons (c) 36 electrons (d) 72 electrons Answer -(b) 10 electrons 10. Given that the ionization enthalpy of hydrogen atom is 1.312 × 106 J mol-1. Then the required amount of energy for the excitation of the electron in the atom from n = 1 to n = 2 is (a) 8.51 × 105 J mol-1 (b) 6.56 × 105 J mol-1 (c) 7.56 × 105 J mol-1 (d) 9.84 × 105 J mol-1 Answer-(d) 9.84 × 105 J mol-1 11. The formula for a mass number of an atom is (a)Number of protons + number of electrons. (b)Number of neutrons + number of electrons. (c)Number of protons + number of neutrons. (d) Non of these Answer -(c)Number of protons + number of neutrons. 12. The atomic number during a chemical reaction will be (a)Increases. (b)Changes. (c)Remains the same. (d)Changes alternatively. Answer -(c)Remains the same. 13. Who proposed the atomic theory? (a)John Dalton. (b)Robert Millikan. (c)J. J. Thomson. (d)Neils Bohr. Answer -(a)John Dalton. 14. The correct electronic configuration of element having atomic number 20 is (a)2, 6, 6, 2 (b)2, 8, 8, 2 (c)2, 4, 6, 2 (d)2, 4, 6, 2 Answer -(b)2, 8, 8, 2 15. An atom differs from an ion as (a)Number of protons. (b)Nuclear charge. (c)Number of electrons. (d)Mass number. Answer -(c)Number of electrons. 16. The place of atom where maximum mass of it concentrated is (a)Nucleus. (b)Neutrons. (C)Protons. (d)Electrons. Answer -(a)Nucleus. 17. CO is iso-electronic with (a) O–2 (b) N+2 (c) CN– (d) O+2 Answer -(c) CN– 18. Which of the following corresponds to the line spectrum of hydrogen obtained in the visible region of light? (a) Lyman series (b) Balmer series (c) Paschen series (d) Brackett series Answer -(b) Balmer series 19. Which of following will be isolates, If the nitrogen atom had electronic configuration $1s^2$, it would have energy lower than that of the normal ground state configuration $1s^2\ 2s^2\ 2p^3$, as the distance between electrons and the nucleus is too less. Yet, $1s^2$ is not observed. (a) Heisenberg’s Uncertainty Principle (b) Hund’s rule (c) Pauli Exclusion Principle (d) Bohr postulate of stationary orbits Answer-(c) Pauli Exclusion Principle 20. if the quantum numbers + 1/2 and – 1/2 for the electron spin it resembles (a) rotation is clockwise and anticlockwise direction respectively (b) rotation is anticlockwise and clockwise direction respectively (c) m Is up and down respectively (d) 2 quantum mechanical spin states which have no classified analogue Answer -(d) 2 quantum mechanical spin states which have no classified analogue 21. What will be the possible energy value of the excited state for electron in Bohr orbitals of hydrogen is , (Given energy of electron in the first Bohr orbit of H atom is 13.6 eV.) (a) – 3.4 eV (b) -4.2eV (c) – 6.8 eV (d) + 6.8 eV Answer-(a) – 3.4 eV 22. What will be the atomic number of the element , if the last entering electron in an element has quantum number n = 3, l = 2, m = + 2 and s = + 1 /2. (a) 13 (b) 21 (c) 29 (d) 39 Answer- (c) 29 23. Given energy = 2.91 × 10-19], h = 6.36 × 10-34 Js, c = 3.0 × 108 m/s ,Then what will be the wavelength of light? (a) 6.56 nm (b) 656 nm (c) 0.656 nm (d) 65.6 nm Answer- (b) 656 nm 24. Which of the following statement is explained by Bohr’s atomic model? (a) the spectrum of hydrogen atom only (b) The spectrum of an atom/ion having only one electron. (c) the solar spectrum Answer- (b) The spectrum of an atom/ion having only one electron. 25. Specification of Magnetic quantum number is (a) orbital size (b) orbital shape (c) orbital orientation (d) nuclear stability Answer- (c) orbital orientation 26. Accurate set of quantum numbers belong to highest energy is (a) n = 4, l = 0, m = 0, s = + 12 (b) n = 3, l = 0, m = 0, s = + 12 (c) n = 2, l = 1, m = 1, s = + 12 (d) n = 3, l = 2, m = 1, s = + 12 Answer- (d) n = 3, l = 2, m = 1, s = + 12 27. Write The correct sequence to label the subshells in an atom is (a)S, p, d, f, g (b)S, p, p, f, d (c)S, s, p, p, d, f, g (d)S, p, g, d, f Answer - (a)S, p, d, f, g 28. Which rule will this configuration be violating,if electronic configuration for oxygen is written as 1s2 2s2 2p4. (a)Aufbau’s principle. (b)Hund’s rule. (c)Pauli’s exclusion principle. (d)None of the above. Answer (d)None of the above. 29. If an atom has four unpaired electrons. Then total spin of electron is (a)1 (b)2.5 (c)2 (d)4 Answer (c)2 30. The charge and mass ratio of cathod rays resembles to (a) a-particles (b) anode rays (c) B-rays (d) protons. Answer (c) B-rays 31.The nucleus of an atom has (a) protons and neutrons. (b) protons and electrons (c) neutrons and electrons. (d) protons, neutrons and electrons. Answer (a) protons and neutrons. 32. Generally Cathode rays are known as (a) protons (b) electrons (c) neutrons (d) a-particles Answer (b) electrons 33. Proton may be defined as (a) an ionised hydrogen molecule (b) an alpha-ray particle (c) a fundamental particle (d) nucleus of heavy hydrogen Answer (c) a fundamental particle 34. The order of size of nucleus ia (a) 10^{-8}cm (b) 10^{-13}cm (c) 10^{-10}cm (d) 10^{-18} cm. Answer (b) 10^{-13}cm 35. The electron having a mass of (a) 9.11×10^{-27} g (b) 9.11×10^{-23} g (c) 9.11 x 10^{-28}g (d) 9.11×10^{-38} g. Answer (c) 9.11 x 10^{-28}g 36. The charge on proton is (a) 1.602 10-19 C (b) 1.502 10-19 C (c) 1.602 10-20 C (d) 1.502 10-21 C. Answer (a) 1.602 10-19 C 37. The average distance of electron from the nucleus within an atom is equals to the order of (a) I cm (b) 10^{13}cm (c) 10^{-8}cm (d) 10^{-12}cm. Answer (c) 10^{-8}cm 38. The founder of anode rays (a) Crookes (b) Goldstein (c) Thomson (d) Chadwick. Answer (b) Goldstein 39.The Canal rays are known as (a) electrons (b) neutrons (c) protons (d) positively charged ions. Answer (d) positively charged ions. 40. Except which of the following, neutrons are present in all atoms? (a) H (b) Ar (c) He (d) Ne Answer (a) H 41. Cl (34,17)and Cl (37,17)differ from each other in number of (a) nucleon (b) positrons (c) protons (d) electrons. Answer (a) nucleon 42. Isotones are those which have same (a) atomic number (b) mass number (c) number of neutrons (d) number of electrons. Answer (c) number of neutrons 43. An isotone of Ge(76,32) is (a) Ge(77,32) (b) Kr (81,36) (c) Se(77,34) (d) As(77,33) Answer (d) As(77,33) 44. The non-integral atomic masses having by many elements is due to (a) they have isobars (b) their isotopes have non-integral masses (c) their isotopes have different masses (d) the constituents, neutrons, protons and electrons. combine to give fractional masses. Answer (c) their isotopes have different masses 45. The number of electrons in an atom is same as its (a) atomic weight (b) atomic number (c) equivalent weight (d) electron affinity. Answer (b) atomic number 46. Neon iso-electronic with (a) O (b) Mg^{2+} (c) N- (d) F^{-2} Answer (b) Mg^{2}+ 47. What is the ratio of charge to mass that was determined by J.J Thomson (in coulombs per gram) (a) -1.76 10^{8}coulombs/ g (b) 1.76 10^{-8} coulombs/g (c) -1.76x 10^{10} coulombs/g (d) -1.76×10^{-10} coulombs /g. Answer (a) -1.76 10^{8}coulombs/ g 48. Rutherford’s scattering experiment explains about (a) the size nucleus (b) size electron (c) radius of atom (d) radius neutron Answer (a) the size nucleus 49. According to Rutherford’s theory, which of the following statement is not true? (a) alpha particles going near the nucleus are slightly deflected (b) some of alpha -particles pass through the nucleus (c) Most of the bita-particles pass through without deflection (d) A few alpha particles are deflected back Answer (b) some of alpha -particles pass through the nucleus 50. The mass of a neutron is (a) same as that of a proton (b) slightly less than proton (c) slightly more than a proton (d) same as that of an electron. Answer (c) slightly more than a proton 51. Rutherford’s experiment of a particles showed for the first time that atom has (a) nucleus (b) electrons (c) protons (d) neutrons Answer (a) nucleus 52. Which established the nuclear model of the atom used a beam of(basis of Rutherford’s theory) (a) bita-particles, which impinged on a metal foil . (b) Bita-rays, which impinged on a metal foil and ejected electron (c) helium atoms, which impinged on a metal foil andst scallered (d) helium nuclei, which impinged on a metal fail and got scallerod Answer (d) helium nuclei, which impinged on a metal fail and got scallerod MCQ on Quantum Numbers 1. For l = 1, the shape of the orbital is (a) Unsymmetrical (b) Spherically symmetrical (c) Dumb-bell (d) Complicated Answer. (c) Dumb-bell 2. The principal quantum number represents (a) Shape of an orbital (b) Distance of electron from nucleus (c) Number of electrons in an orbit (d) Number of orbitals in an orbit Answer. (b) Distance of electron from nucleus 3. How many electrons can be fit into the orbitals that comprise the 3rd energy level ? (a) 2 (b) 8 (c) 18 (d) 32 Answer. (c) 18 4. If value of l = 0, the shape of the orbital is (a) Rectangular (b) Spherical (c) Dumbbell (d) Unsymmetrical Answer. (b) Spherical 5. For n = 3, the number of possible orbitals are (a) 1 (b) 3 (c) 4 (d) 9 Answer. (d) 9 6.When 3d orbital is complete, the new electron will enter the (a) 4p-orbital (b) 4f-orbital (c) 4s-orbital (d) 4d-orbital Answer. (a) 4p-orbital 7. Total number of radial nodes of 3s and 2p orbitals are respectively. (a) 2, 0 (b) 0, 2 (c) 1, 2 (d) 2, 1 Answer. (a) 2, 0 8. The mvr of an electron depends on (a) Principal quantum number (b) Azimuthal quantum number (c) Magnetic quantum number (d) All of these Answer. (b) Azimuthal quantum number 9. The maximum energy is present in any electron at which area (a) Nucleus (b) Ground state (c) First excited state (d) Infinite distance from the nucleus Answer. (d) Infinite distance from the nucleus 10.The number of orbitals for n = 4 will be (a) 4 (b) 8 (c) 12 (d) 16 Answer. (d) 16 11.The two electrons in K sub-shell will differ in (a) n (b) l (c) m (d) s Answer. (d) s 12.The element having atomic number 29 possess how many unpaired electrons in d-orbitals (a) 10 (b) 1 (c) 0 (d) 5 Answer. (c) 0 13. The similarities found between 2p and 3p orbital that is (a) Shape (b) Size (c) Energy (d) Value of n Answer. (a) Shape 14. The azimuthal quantum number is related to (a) Size (b) Shape (c) Orientation (d) Spin Answer. (b) Shape 15.Value of l for last electron of Na atom is (a) 1 (b) 2 (c) 3 (d) 0 Answer. (d) 0 16. In d orbitals the maximum number of unpaired electron can be present is (a) 1 (b) 3 (c) 5 (d) 7 Answer. (c) 5 17. The number of unpaired electrons in an O2 molecule is (a) 0 (b) 1 (c) 2 (d) 3 Answer. (c) 2 18.The number of unpaired electron I the element for Z=29 (a) 1 (b) 3 (c) 4 (d) 2 Answer. (a) 1 19. Maximum how many quantum numbers may same by two electrons (a) One (b) Two (c) Three (d) Four Answer. (c) Three 20. How many orbitals can have the following quantum numbers, n = 3, l = 1, ml = 0? (a) 4 (b) 2 (c) 1 (d) 3 Answer: (c) 1 21. For the set of quantum numbers, n = 3, l = 1, m = -1, the maximum electrons will be (a) 2 (b) 6 (c) 10 (d) 4 Answer: (a) 2 22. The maximum number of electrons that can fit in an orbital with n = 3 and l = 1? (a) 14 (b) 6 (c) 10 (d) 2 Answer: (d) 2 23. Which of the following helps to determine the maximum number of electrons present in a subshell (a) 2l + 1 (b) 2n2 (c) 4l + 2 (d) 4l – 2 Answer: (c) 4l + 2 24. Which of the following quantum numbers governs the spatial orientation of an atomic orbital? (a) Magnetic quantum number (b) Spin quantum number (c) Azimuthal quantum number (d) Principal quantum number Answer: (a) Magnetic quantum number 25. Which one represents the three-dimensional shape of an atomic orbital (a) Azimuthal quantum number (b) Principal quantum number (c) Spin quantum number (d) Magnetic quantum number Answer: (a) Azimuthal quantum number 26. The maximum number of orbitals present in a subshell can be represented by (a) 2l + 1 (b) 2n2 (c) 4l + 2 (d) 4l – 2 Answer: (a) 2l + 1 27. For the orbitals with n = 2 and l = 1, the number of electrons that fit in it is (a) 8 (b) 2 (c) 6 (d) 4 Answer: (c) 6 28. For l=3 the number of electrons will be (a) 14 (b) 2 (c) 10 (d) 6 Answer: (a) 14 Also You Can Read Inorganic Chemistry MCQ

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